

POSITIVE IONS (CATIONS)	NEGATIVE IONS (ANIONS)
GROUP 1 Elements 1+ examples: lithium Li^+ sodium Na^+ potassium K^+	GROUP 16 Elements 2- the names end in "ide" examples: oxide O^{2-} sulfide S^{2-}
GROUP 2 Elements 2+ examples: beryllium Be^{2+} magnesium Mg^{2+} calcium Ca^{2+}	GROUP 17 Elements 1- the names end in "ide" examples: fluoride F^- chloride Cl^-
POLYATOMIC CATION ammonium NH_4^+	POLYATOMIC ANIONS acetate $\text{C}_2\text{H}_3\text{O}_2^-$
TRANSITION ELEMENTS require Roman numerals copper(I) Cu^+ copper(II) Cu^{2+}	carbonate CO_3^{2-} hydrogen carbonate (bicarbonate) HCO_3^-
iron(II) Fe^{2+} iron(III) Fe^{3+}	chlorate ClO_3^- chromate CrO_4^{2-}
lead(II) Pb^{2+} lead(IV) Pb^{4+}	hydroxide OH^-
mercury(I) Hg_2^{2+} mercury(II) Hg^{2+}	nitrate NO_3^- nitrite NO_2^-
nickel(II) Ni^{2+}	oxalate $\text{C}_2\text{O}_4^{2-}$
tin(II) Sn^{2+} tin(IV) Sn^{4+}	permanganate MnO_4^-
<i>(not requiring Roman numerals)</i> zinc Zn^{2+} cadmium Cd^{2+} silver Ag^+	phosphate PO_4^{3-}
MISC aluminum Al^{3+} hydrogen H^+	sulfate SO_4^{2-} sulfite SO_3^{2-}
	hydride H^- nitride N^{3-}